George Mason University
CHEM 321: 15 MINUTE QUIZ
Chapters 1 and 3 **Name _____**

(Show work to get full credit)

1. (5 points)
How many grams of solid KF should be added to 10.0 L of water to give 1.2 ppm F^-?
(At wts: F = 18.998, K = 39.098)

2. (5 points)
Given the following equilibria:
$Ag^+ + Cl^- \Leftrightarrow AgCl\ (aq)$ K = 2000.0
$AgCl\ (s) \Leftrightarrow Ag^+ + Cl^-$ $K = 1.8 \times 10^{-10}$

Find K for $AgCl\ (aq) \Leftrightarrow AgCl\ (s)$.

3. (10 points)
To prepare a solution of NaCl, you weigh out 2.567 (± 0.002) g and dissolve it in a volumetric flask of 100.00 (±0.10) mL in volume. Express the molarity of the solution and its uncertainty with appropriate number of significant figures. (Formula mass of NaCl is 58.443 ±0.002 g/mol)